

# Experiment 8

## Limiting Reactant

Colcium avalate crystals contribute to the formation of kidney stones.

- To determine the limiting reactant in a mixture of two soluble salts
- To determine the percent composition of each substance in a salt mixture

#### OBJECTIVES

The following techniques are used in the Experimental Procedure:













TECHNIQUES

Percent composition: the mass ratio of a component of a mixture or compound to the total mass of the sample times 100









Two factors affect the yield of products in a chemical reaction: (1) the amounts (moles) of starting materials (reactants) and (2) the percent yield of the reaction. Many experimental conditions, for example, temperature and pressure, can be adjusted to increase the yield of a desired product in a chemical reaction, but because chemicals react according to fixed mole ratios (stoichiometrically), only a limited amount of product can form from measured amounts of starting materials. The reactant determining the amount of product generated in a chemical reaction is called the **limiting reactant** in the chemical system.

To better understand the concept of the limiting reactant, let us look at the reaction under investigation in this experiment, the reaction of calcium chloride dihydrate, CaCl<sub>2</sub>·2H<sub>2</sub>O, and potassium oxalate monohydrate, K<sub>2</sub>C<sub>2</sub>O<sub>4</sub>·H<sub>2</sub>O, in an aqueous solution.

$$CaCl_2 \cdot 2H_2O(aq) + K_2C_2O_3 \cdot H_2O(aq) \rightarrow$$
  
 $CaCl_2O_4 \cdot H_2O(s) + 2 \text{ KCl}(aq) + 2 \text{ H}_2O(l)$  (8.1)

Calcium oxalate monohydrate, CaC<sub>2</sub>O<sub>2</sub>H<sub>2</sub>O, the product, is an insoluble compound, but is found naturally in a number of diverse locations. It is found in plants, such as rhubarb leaves, agave, and (in small amounts) spinach, and is the cause of most kidney stones. In small doses, it causes a severe reaction to the lining of the digestive tract. However, the handling of calcium oxalate in the laboratory is safe, so long as it is not transferred to the mouth.

For the reaction system in this experiment, both the calcium chloride and potassium oxalate are soluble salts, but the calcium oxalate is insoluble. The **ionic equation** for the reaction is

$$Ca^{2*}(aq) + 2 C\Gamma(aq) + 2 K^*(aq) + C_2O_4^{-2}(aq) + 3 H_2O(l) \rightarrow CaC_2O_4 \cdot H_2O(s) + 2 C\Gamma(aq) + 2 K^*(aq) + 2 H_2O(l)$$
 (8.2)

#### INTRODUCTION

Percent yield:

Stoichiometrically: by a study of a chemical reaction using a balanced equation

lonic equation: A chemical equation that presents ionic compounds in the form in which they exist in aqueous solution. See Experiment 6.

## **Experiment 8 Limiting Reactant Answers**

Mary Millhollon, Richard H. Langley

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