



Final Discussion Guide

Objective	<p>This guide will help you, the support person, lead the final conversation with a new [redacted]. Use your knowledge and experience to help answer participants questions and to gauge their understanding of the topics covered in the course. The following is a suggested approach with sample conversation questions.</p> <p>This session takes approximately 30 minutes.</p>
Preparation	<p>Do the following before the session starts:</p> <ul style="list-style-type: none"> • Complete all three training courses. • Review the learning objectives for each course. • Review the answer keys for each course, specifically the scenarios. • [redacted] should be reminded to bring their completed workbooks to this session (four in total).
Ask	<p>Ask the [redacted] about their experience completing the training.</p> <ul style="list-style-type: none"> • Discuss their experience and clarify, as needed. • Encourage them to share by asking about specific parts of the training courses.
Ask	<p>Ask the [redacted] what general questions they have about the topics covered in the training.</p> <ul style="list-style-type: none"> • This is a precursor to reviewing the scenarios and discussing the [redacted] answers to the questions.
Discuss Scenario #1	<p>Ask [redacted] to explain their understanding of the first scenario. You can use the following questions and the answer key, as needed. The answer key is available at the location noted in the Answer Keys section of this document.</p> <ul style="list-style-type: none"> • What happened in this situation? • How would you respond? • What follow-up activities would you recommend in this instance? • What questions did you have about this scenario?
Discuss Scenario #2	<p>Ask [redacted] to explain their understanding of scenario #2. You can use the questions from scenario #1 and the answer key, as needed.</p>
Discuss Scenario #3	<p>Ask [redacted] to explain their understanding of scenario #3.</p>
Discuss Scenario #4	<p>Ask [redacted] to explain their understanding of scenario #4.</p> <ul style="list-style-type: none"> • Ask how their recommended response/follow-up might change if instead of [redacted], Matt had told his [redacted] that he was [redacted].
Ask	<p>Ask the [redacted] if they have other questions or any aspects of [redacted] that they would like to review.</p>
Wrap-Up	<p>Inform the [redacted] that the session is complete.</p> <p>Let them know that they can come back to you if they have further questions or concerns.</p>

Entwined With You Discussion Guide

Janet E. Salmons



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web there are three primary theories of acid base chemistry that are often taught together arrhenius theory brønsted lowry theory and lewis acid base theory each theory is introduced below figure 2 3 1 2 3 1 hierarchical definitions of acids and bases via the three primary theories these theories are designed to be superset of the previous

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to include those reactions where no proton transfer takes place according to this concept an acid

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web arrhenius bases an arrhenius base is defined as any species that increases the concentration of hydroxide ions OH^- in aqueous solution an example of an arrhenius base is the highly soluble sodium hydroxide NaOH sodium hydroxide dissociates in water as follows $\text{NaOH} \rightarrow \text{Na}^+ + \text{OH}^-$ in water sodium hydroxide

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web work done by svante arrhenius in the late 19th century set the groundwork for our current understanding of acid base theory despite the relatively limited technology available to him he was able to develop theories that adequately explained the properties and chemical reactivity of many acids and bases

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web the arrhenius theory of acids and bases acids are substances that produce hydrogen ions in solution bases are substances that produce hydroxide ions in solution neutralization happens because hydrogen ions and hydroxide ions react to produce water $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$ $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$ hydrochloric

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web aug 15 2020 the equilibrium formulated in scheme 1 between hydrogen ion and the corresponding acid and base can be called a simple acid base equilibrium by mixing two simple systems a double acid base system and an acid base equilibrium result that can always be formulated as follows $\text{acid}_1 + \text{base}_2 \rightleftharpoons \text{acid}_2 + \text{base}_1$

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web key points a brønsted lowry acid is any species that is capable of donating a proton H^+ a brønsted lowry base is any species that is capable of accepting a proton which requires a lone pair of electrons to bond to the H^+ water is amphoteric

which means it can act as both a brønsted lowry acid and a brønsted lowry base

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web the a ion is a base because it is accepting a proton from the h 3 o the reversible reaction contains two acids and two bases we think of them in pairs called conjugate pairs when the acid ha loses a proton it forms a base a when the base a accepts a proton back again it obviously reforms the acid ha

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