

## K<sub>sp</sub>

1. **Solubility Equilibrium**-The heterogeneous equilibrium involved in the dissolution or precipitation of ionic compounds. This will tell us the amount of a substance that will dissolve (solubility).

2. **Vocabulary**-

- a. saturated solution- a solution that contains undissolved solute in contact with the solvent.
- b. ionic solid- totally ionizes into its constituents.

3. **Solubility Product Constant** – aka solubility product, K<sub>sp</sub>, indicates how soluble the solid is in water.

a. Given an ionic compound BA, there is an equilibrium condition that occurs even though the compound completely ionizes. At the saturation point, ions are reforming the solid at the same rate that the solid is forming ions:



b. Remember:

Solubility Constant –

has one value for a given solid at a given temperature.

is the molar concentration of ions raised to their stoichiometric powers.

Solubility – a equilibrium position.

Solubility changes as concentrations change.

Solubility is the amount (grams) of substance that dissolves to form a saturated solution

Molar Solubility- a new unit

Molar solubility is the number of moles of solute dissolving to form a liter of saturated solution.

4. **Conversion between the units**

a. To convert solubility to K<sub>sp</sub>

solubility needs to be converted into molar solubility (via molar mass);

molar solubility is converted into the molar concentration of ions at equilibrium (equilibrium calculation).

K<sub>sp</sub> is the product of equilibrium concentration of ions.



### Calculating K<sub>sp</sub> from Solubility I

Copper(I) bromide has a measured solubility of  $2.0 \times 10^{-4}$  mol/L at 25°C. Calculate its K<sub>sp</sub> value.

# Equilibrium And Ksp Study Guide

**Robert K. Wismer**



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