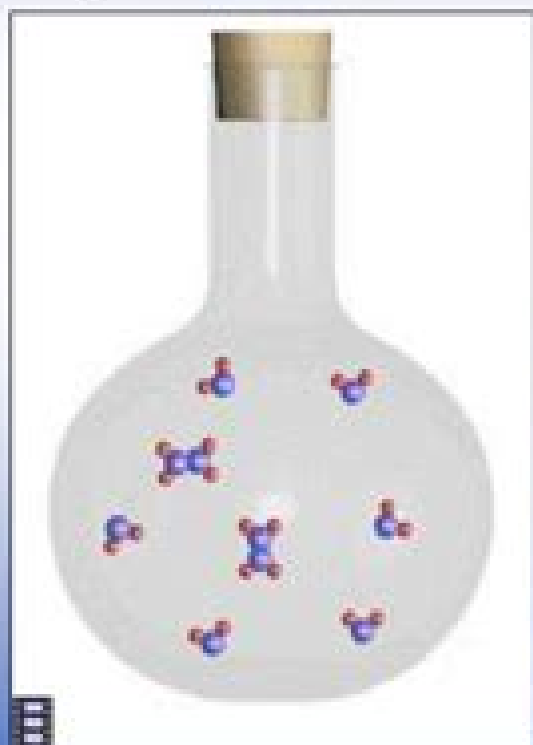


Nitrogen Dioxide Equilibrium

Increase P in the system by reducing the volume (at constant T).



$$K = \frac{[\text{NO}_2]^2}{[\text{N}_2\text{O}_4]} = 0.0058 \text{ at } 298 \text{ K}$$

In gaseous system the equilibrium will shift to the side with **fewer molecules** (in order to reduce the P).

Therefore, reaction shifts _____
and P of NO_2 _____ and P of
 N_2O_4 _____.

Equilibrium Nitrogen Dioxide Lab

Darryl A. Wahlstrom

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Equilibrium Nitrogen Dioxide Lab:

Technical Report - Jet Propulsion Laboratory, California Institute of Technology Jet Propulsion Laboratory (U.S.),1962
Publications of the Jet Propulsion Laboratory, January 1938 Through June 1961 D. K. Walsh,1961 *Laboratory Exercises in Inorganic Chemistry* James Flack Norris,Kenneth Lamartine Mark,1922 **Lecture and Laboratory Experiments in Physical Chemistry** Henry Sjoerd van Klooster,1925 *The Dinitrogen Tetroxide/nitrogen Dioxide Gas-phase Equilibrium* Darryl A. Wahlstrom,1989 The Heat Capacity of the System Nitrogen Tetroxide and Nitrogen Dioxide Edgar DeBolt McCollum,1925 Introduction to Chemistry, Laboratory Manual T. R. Dickson,1994-12-23 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and a chemical vision learning problem solving methods and utilizing group work and discussion activities These themes involve and engage the students in their own learning processes they are challenged to be active The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics The reorganization presents many core concepts in the first five chapters preparing students for later chapters In addition the author has added vignettes throughout the chapters referring to health technology the environment and society as well as to specific tools of direct use to students **Laboratory manual** William McPherson,1913 A Laboratory course in physiology Walter Bradford Cannon,1911 **Laboratory Manual, Arranged to Accompany the 3d Ed. of "A Course in General Chemistry"** William McPherson,William Edwards Henderson,William Lloyd Evans,1928 **A Laboratory Manual of General Chemistry for Use in Colleges** William Crowell Bray,Wendell Mitchell Latimer,1921 **EPA Publications Bibliography Quarterly Abstract Bulletin** United States. Environmental Protection Agency,1992 **The Atmospheric Diagnostics Program at Brookhaven National Laboratory** Walter D. Tucker,1969 Laboratory Experiments in General Chemistry George Brooks King,William Elmer Caldwell,1967 Contribution from the Research Laboratory of Physical Chemistry Massachusetts Institute of Technology. Research Laboratory of Physical Chemistry,1920 Physiology by the Laboratory Method William Joshua Brinckley,1902 **A Laboratory Manual of General Chemistry** William Crowell Bray,Ludwig Rosenstein,1916 **Laboratory Directions and Study Questions in Inorganic Chemistry** Alexander Silverman,1919 **Laboratory Manual for General Chemistry** Alvin Strickler,1923 **Journal of the Society of Chemical Industry** Society of Chemical Industry (Great Britain),1927 Includes list of members 1882 1902 and proceedings of the annual meetings and various supplements

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