

Name: \_\_\_\_\_ Student Number: \_\_\_\_\_ Date: \_\_\_\_\_

## Student Exploration: Collision Theory

**Vocabulary:** activated complex, catalyst, chemical reaction, concentration, chemical formula, molecular, product, reactant, surface area

**Prior Knowledge Questions** (Do this before starting the simulation.)

1. Suppose you added an equal amount of sugar to two glasses and poured the two glasses of water. Would you expect them to dissolve completely? \_\_\_\_\_ How fast would? \_\_\_\_\_
2. Suppose you took a liquid mixture and added more of one of the liquid and another made the mix of more of the other. Would the mixture of more of the other be more likely? \_\_\_\_\_ How fast would? \_\_\_\_\_

**Before the simulation**

In **collision theory**, a reaction that does not occur spontaneously is **activated**. **Activation energy** is the minimum amount of energy that **reactants** must have to overcome the activation energy barrier. **Reactants** are substances that participate in the reaction. The **activation energy** barrier is the energy that must be overcome for a reaction to occur. The **activated complex** is the intermediate state between the reactants and the products in a chemical reaction.

The **activated complex** is the state of the reaction where the reactants are in the process of becoming products. It is the highest energy state in the reaction.



3. Look at the ball in the picture above. Which of the spheres is the green sphere? Which is the blue sphere? Which is the yellow sphere? Which is the red sphere?

Reactants:

Products:

4. How many **collisions** occur in the reaction? \_\_\_\_\_ How many of the collisions result in a reaction? \_\_\_\_\_ How many of the collisions result in a reaction? \_\_\_\_\_

# Exploration Guide Collision Theory Gizmo Answer Key

**Huangqi Zhang**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the author's name.

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