

### Recap/Summary

1. Study of this \_\_\_\_\_ method \_\_\_\_\_ relationships in a \_\_\_\_\_ situation \_\_\_\_\_
2. Research on \_\_\_\_\_ behaviour \_\_\_\_\_ equated

2.  $40g + 40g \rightarrow 2 \text{ MgO}$

2. Magnesium burning with oxygen, oxygen being provided in formula with  
at  
magnesium oxide

The \_\_\_\_\_ in a \_\_\_\_\_ reaction \_\_\_\_\_ give the  
\_\_\_\_\_ \_\_\_\_\_ from the \_\_\_\_\_ involved in the  
reaction \_\_\_\_\_

E.g. Problem

When aluminium aluminium reacts with aluminium oxide, aluminium chloride is  
produced



moles reaction



2. If you start with 4 moles of Al, how many moles of  $\text{Al}_2\text{O}_3$  will be produced?

2.  $2 \text{ Al} \rightarrow \text{Al}_2\text{O}_3$  and  $3 \text{ O}_2 \rightarrow 2 \text{ Al}_2\text{O}_3$

# Episode 801 Note Taking Guide Key

**RC Schank**



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