

Experiment 23

Data and Calculations: Determination of the Equilibrium Constant for a Chemical Reaction

Mixture	Volume in mL, 2.00×10^{-3} M $\text{Fe}(\text{NO}_3)_3$	Volume in mL, 2.00×10^{-3} M KSCN	Volume in mL, Water	Method I Absorbance	Method II Depth in mm		$[\text{FeSCN}^{2+}]$
					Standard	Unknown	
1	5.00	1.00	<u>4.00</u>	<u>0.084</u>	—	—	<u>0.25×10^{-4} M</u>
2	5.00	2.00	<u>3.00</u>	<u>0.218</u>	—	—	<u>0.45×10^{-4} M</u>
3	5.00	3.00	<u>2.00</u>	<u>0.317</u>	—	—	<u>0.60×10^{-4} M</u>
4	5.00	4.00	<u>1.00</u>	<u>0.447</u>	—	—	<u>0.84×10^{-4} M</u>
5	5.00	5.00	<u>0</u>	<u>0.514</u>	—	—	<u>1.00×10^{-4} M</u>

If Method II was used, $[\text{FeSCN}^{2+}]_{\text{unknown}} = \text{—} \times 10^{-4}$ M; $[\text{FeSCN}^{2+}]$ in Mixtures 1 to 5 is found by Equation 7.

Processing the Data

A. Calculation of K_c assuming the reaction:



This calculation is most easily done by following Steps 1 through 5 in the discussion. Results are to be entered in the table on the following page. If you are using Excel, set up the table as we have, and follow the directions below.

Step 1 Find the initial number of moles of Fe^{3+} and SCN^{-} in the mixtures in test tubes 1 through 5. Use Equation 3 and enter the values in the first two columns of the table.

Step 2 Enter the experimentally determined value of $[\text{FeSCN}^{2+}]$ at equilibrium for each of the mixtures in the next to last column in the table. Use Equation 3 to find the number of moles of FeSCN^{2+} in each of the mixtures, and enter the values in the fifth column of the table. Note that this is also the number of moles of Fe^{3+} and SCN^{-} that were used up in the reaction.

Experiment 23 Determination Equilibrium Constant

Dr. M. Satish Kumar



Experiment 23 Determination Equilibrium Constant:

Chemistry in the Laboratory James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations A Course of Laboratory Experiments on Physico-chemical Principles Miles Standish Sherrill, 1923 **Experimental Chemistry** Russell A. Geanangel, Wesley William Wendlandt, 1979-08 **Experiments in General Chemistry** Gerald S. Weiss, 1985

Experimental Physical Chemistry Arthur Halpern, George McBane, 2006-06-30 Experimental Physical Chemistry includes complete lists of necessary materials detailed background material for each experiment and relevant sections on measurements and error analysis *General Chemistry Laboratory Text with Qualitative Analysis* Wilbert Hutton, 1968

Nuclear Science Abstracts, 1969 NSA is a comprehensive collection of international nuclear science and technology literature for the period 1948 through 1976 pre dating the prestigious INIS database which began in 1970 NSA existed as a printed product Volumes 1 33 initially created by DOE s predecessor the U S Atomic Energy Commission AEC NSA includes citations to scientific and technical reports from the AEC the U S Energy Research and Development Administration and its contractors plus other agencies and international organizations universities and industrial and research organizations References to books conference proceedings papers patents dissertations engineering drawings and journal articles from worldwide sources are also included Abstracts and full text are provided if available **Chemical Principles in the Laboratory** Emil J. Slowinski, 1996 Provides a series of experiments designed to teach students the available experimental methods the proper design of experiments and the interpretation of experimental results *Experiments in General Chemistry* Kenneth W. Whitten, Carl B. Bishop, Kenneth D. Gailey, 1992 **Laboratory Manual** Jo A. Beran, 1990

Chemical Principles in Practice Jerry A. Bell, 1967 This forward looking laboratory text includes thirty three experiments designed to serve high level introductory chemistry courses It is intended in part to give the students practice in techniques not only those generally considered the techniques of chemistry but also those of mathematics However the only prerequisite is a good background in high school mathematics Ideally suited for use with University Chemistry or College Chemistry by Bruce Mahan this book follows the physical chemistry approach of these and most other modern chemistry texts maintaining the same flavor emphasis and degree of sophistication All experiments are open ended and include suggestions for following them up There is considerable emphasis on the involvement of the student in planning his own procedures and quantitative thinking Publisher Laboratory Experiments in Chemistry J. J. Lagowski, Stephen E. Webber, 1977 *Experiments in Physics, for Students of Science and Engineering* Ernest Blaker, 1905 Practical Physical Chemistry Dr. M. Satish Kumar, Physical Chemistry deals with the relations between the physical properties of substances and their composition The present

book is intended to serve as a practical manual for undergraduate and post graduate students I have attempted to assemble the list of experiments from my experience and also have drawn upon the experience of the students who have undergone these laboratory courses and felt the inadequacy of the existing syllabus I am aware that I have not yet exhausted all the experiments that they wanted to place in this book but I had to make a selection keeping the size in consideration This manual is largely structured around the standard experiments of physical chemistry Detailed information on instrumentation kinetics experimental methods and data analysis has been covered I will be happier to take all comments and incorporate them in the further editions

Experimental Chemistry Richard C. Hatch,1972 *Biothermodynamics, Part B* Michael L. Johnson,Jo M. Holt,Gary K. Ackers,2009-11-11 The use of thermodynamics in biological research can be equated to an energy book keeping system While the structure and function of a molecule is important it is equally important to know what drives the energy force These methods look to answer What are the sources of energy that drive the function Which of the pathways are of biological significance As the base of macromolecular structures continues to expand through powerful techniques of molecular biology such as X ray crystal data and spectroscopy methods the importance of tested and reliable methods for answering these questions will continue to expand as well This volume presents sophisticated methods for estimating the thermodynamic parameters of specific protein protein protein DNA and small molecule interactions Elucidates the relationships between structure and energetics and their applications to molecular design aiding researchers in the design of medically important molecules Provides a must have methods volume that keeps MIE buyers and online subscribers up to date with the latest research Offers step by step lab instructions including necessary equipment from a global research community

General Chemistry Wilbert Hutton,1965 **Chemistry in the Laboratory** John J. Alexander,Margaret J. Steffel,1976 **For the Implementation of Science Syllabus, Years 11 and 12** New South Wales. Department of Education,1980 SUMMARY Suggestions to assist teachers with the implementation of the 2 Unit Course in Chemistry

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