

EXPERIMENT: INVESTIGATING PRECIPITATION REACTIONS

PDF

AIM

To compare the solubility of a range of ionic substances through precipitation.

EQUIPMENT

☐ test tubes (6)

Cations (all 0.1 mol L⁻¹) in dropper bottles:

- ☐ sodium nitrate, NaNO₃
- ☐ silver nitrate, AgNO₃
- ☐ lead(II) nitrate, Pb(NO₃)₂
- ☐ ammonium nitrate, NH₄NO₃
- ☐ magnesium sulfate, MgSO₄
- ☐ iron(II) sulfate, FeSO₄
- ☐ copper(II) sulfate, CuSO₄
- ☐ zinc(II) sulfate, ZnSO₄
- ☐ aluminium sulfate, Al₂(SO₄)₃
- ☐ barium chloride, BaCl₂
- ☐ calcium chloride, CaCl₂
- ☐ iron(III) chloride, FeCl₃

☐ test tube rack

Anions (all 0.1 mol L⁻¹) in dropper bottles:

- ☐ potassium nitrate, KNO₃
- ☐ potassium chloride, KCl
- ☐ sodium sulfate, Na₂SO₄
- ☐ sodium carbonate, Na₂CO₃
- ☐ sodium acetate, NaCH₃COO
- ☐ sodium hydroxide, NaOH



SAFETY Wear safety glasses. Lead and barium salts are poisons so avoid contact with skin. If contact occurs wash thoroughly with soap and water. Dispose of as directed by your teacher. Consult "Chemical Safety in Schools" booklet (CSES).

PROCEDURE

- 1 Put a 5 mL of a chemical containing each anion into each of six test tubes. Take care not to mix droppers.
- 2 Add 5 mL of a chemical containing the first cation to each. Record your observations.
- 3 Clean all test tubes thoroughly as incorrect results can arise from contamination of solutions.

Experiment 17 Precipitation Reactions Answers

Léo Pomar



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Table of Contents Experiment 17 Precipitation Reactions Answers

1. Understanding the eBook Experiment 17 Precipitation Reactions Answers
 - The Rise of Digital Reading Experiment 17 Precipitation Reactions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Experiment 17 Precipitation Reactions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Experiment 17 Precipitation Reactions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Experiment 17 Precipitation Reactions Answers
 - Personalized Recommendations
 - Experiment 17 Precipitation Reactions Answers User Reviews and Ratings
 - Experiment 17 Precipitation Reactions Answers and Bestseller Lists
5. Accessing Experiment 17 Precipitation Reactions Answers Free and Paid eBooks
 - Experiment 17 Precipitation Reactions Answers Public Domain eBooks
 - Experiment 17 Precipitation Reactions Answers eBook Subscription Services
 - Experiment 17 Precipitation Reactions Answers Budget-Friendly Options
6. Navigating Experiment 17 Precipitation Reactions Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Experiment 17 Precipitation Reactions Answers Compatibility with Devices
 - Experiment 17 Precipitation Reactions Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Experiment 17 Precipitation Reactions Answers
 - Highlighting and Note-Taking Experiment 17 Precipitation Reactions Answers
 - Interactive Elements Experiment 17 Precipitation Reactions Answers

8. Staying Engaged with Experiment 17 Precipitation Reactions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Experiment 17 Precipitation Reactions Answers
9. Balancing eBooks and Physical Books Experiment 17 Precipitation Reactions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Experiment 17 Precipitation Reactions Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Experiment 17 Precipitation Reactions Answers
 - Setting Reading Goals Experiment 17 Precipitation Reactions Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Experiment 17 Precipitation Reactions Answers
 - Fact-Checking eBook Content of Experiment 17 Precipitation Reactions Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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