EXPERIMENT 19 ACIDIC AND BASIC SALTS

INTRODUCTION

Water makes up about 60% of the weight of a human body; other animals contain a similar proportion of water, and some plants are almost 95% water. Water covers about 70% of the earth's surface; also, it is in the soil and is present as vapor in the atmosphere. Consequently, reactions that occur in aqueous solutions and reactions in which water participates are very important to us; many are essential to the life and growth of organisms. These reactions might affect the pH of water by using or producing hydronium or hydroxide ions.

It is apparent from the formulas of some compounds that they will produce H₂O* or OH* in aqueous solution. An Arrhenius acid contains hydrogen, which is usually written first in the formula to emphasize its ability to react with water producing H₂O*. An ionic Arrhenius base contains hydroxide ions, which separate from the positive ions when the base dissolves in water.

Salts are ionic compounds that contain a cation from a base and an anion from an acid. Even though most salts do not contain H or OH, many of them do react with water changing its pH. Such a reaction is called the hydrolysis of a salt, the reaction of an ion of a dissolved salt with water producing HyO* or OH*.

The hydrolysis of a negative ion produces the negative ion of water. This reaction occurs if the anion is the conjugate base of a weak acid. For example, HCN is a weak acid; $K_a = 4.9 \times 10^{-10}$:

$$HCN + H_2O \Leftrightarrow CN^- + H_3O^*$$
(19-1)

The acid molecule HCN has little tendency to donate protons; its conjugate base, CN*, is a good proton acceptor. Consequently, when a salt containing CN*, such as NaCN, dissolves in water, a Brensted-Lowry acid-base reaction occurs between H₂O (acid) and CN* (base); that is, a proton is transferred from H₂O to CN*:

$$CN^- + H_0O \Leftrightarrow HCN + OH^-$$
 (19-2)

The resulting solution is basic, because hydroxide ions in excess of those produced by the autoionization of water are present.

The hydrolysis of a positive ion produces the positive ion of water. This reaction occurs if the cation is the conjugate acid of a weak base. Molecular bases such as NH₃ are weak bases:

$$NH_b + H_bO \Leftrightarrow NH_b^* + OH^-$$
 (19-3)

For this reaction, $K_b = 1.8 \times 10^{-8}$. The ammonium ion can donate H* to a chemical species that is capable of acting as a base. Consequently, when a salt containing NH₄* dissolves in water, NH₄* and H₂O react in a Brensted-Lowry fashion:

$$NH_4^+ + H_2O \iff NH_3 + H_3O^+$$
 (19-4)

Some hydrated metal ions also hydrolyze producing acidic solutions. The strong attraction between a small, highly charged, metal ion and the oxygen atom of a bound water molecule weakens an oxygen-hydrogen bond in the bound water molecule. As a result, the hydrate acts as a proton donor:

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John Bernard Ekeley

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An Elementary Experimental Chemistry John Bernard Ekeley, 1900 Laboratory Experiments in Chemistry to Accompany Black and Conant's "Practical Chemistry" Newton Henry Black,1920 **Laboratory Experiments in** Experiments and Exercises in Basic Chemistry Steven Murov, Brian **Chemistry** Newton Henry Black, 1921 Stediee.2000-02-03 **Practical Physical Chemistry** Dr. M. Satish Kumar, Physical Chemistry deals with the relations between the physical properties of substances and their composition. The present book is intended to serve as a practical manual for undergraduate and post graduate students I have attempted to assemble the list of experiments from my experience and also have drawn upon the experience of the students who have undergone these laboratory courses and felt the inadequacy of the existing syllabus I am aware that I have not yet exhausted all the experiments that they wanted to place in this book but I had to make a selection keeping the size in consideration This manual is largely structured around the standard experiments of physical chemistry Detailed information on instrumentation kinetics experimental methods and data analysis has been covered I will be happier to take all comments and incorporate them in the further editions **Laboratory Study of Household Chemistry** Mary Ethel Jones, 1921 An Introduction to the study of the compounds of carbon, or, Organic chemistry Ira Remsen, 1894 An Introduction to the Study of the Compounds of Carbon Ira Remsen, 1899 An introduction to the study of the compounds of carbon; or, Organic chemistry. **4th revision** Ira Remsen.1903 Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques. The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures Chemistry Made Simple John T. Moore, Ed.D., 2010-04-21 See the world one molecule at a time Chemistry helps us understand not only the world around us but also our own bodies CHEMISTRY MADE SIMPLE makes it fun Each chapter has practice problems with complete solutions that reinforce learning A glossary of chemical terms the modern periodic table and detailed illustrations throughout make this the best introduction to one of the most studied of all sciences Topics covered include the Scientific Method the structure and properties of matter compounds laws of chemistry gases liquids and solids solutions electrochemistry the atmosphere biochemistry organic chemistry nuclear chemistry energy the environment Look for these Made Simple titles Accounting Made Simple Arithmetic Made Simple Astronomy Made Simple Biology Made Simple Bookkeeping Made Simple Business Letters Made Simple Earth Science Made Simple English Made Simple French Made Simple German Made Simple Ingles Hecho Facil Investing Made Simple Italian Made Simple Latin Made Simple Learning English Made Simple Mathematics Made Simple The Perfect Business Plan Made Simple Philosophy Made Simple Physics Made Simple Psychology Made Simple Sign Language Made Simple Spelling Made

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Table of Contents Experiment 19 Acids And Basic Salts

- 1. Understanding the eBook Experiment 19 Acids And Basic Salts
 - The Rise of Digital Reading Experiment 19 Acids And Basic Salts
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Experiment 19 Acids And Basic Salts
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Experiment 19 Acids And Basic Salts
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Experiment 19 Acids And Basic Salts

- Personalized Recommendations
- Experiment 19 Acids And Basic Salts User Reviews and Ratings
- Experiment 19 Acids And Basic Salts and Bestseller Lists
- 5. Accessing Experiment 19 Acids And Basic Salts Free and Paid eBooks
 - Experiment 19 Acids And Basic Salts Public Domain eBooks
 - Experiment 19 Acids And Basic Salts eBook Subscription Services
 - Experiment 19 Acids And Basic Salts Budget-Friendly Options
- 6. Navigating Experiment 19 Acids And Basic Salts eBook Formats
 - o ePub, PDF, MOBI, and More
 - Experiment 19 Acids And Basic Salts Compatibility with Devices
 - Experiment 19 Acids And Basic Salts Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Experiment 19 Acids And Basic Salts
 - Highlighting and Note-Taking Experiment 19 Acids And Basic Salts
 - Interactive Elements Experiment 19 Acids And Basic Salts
- 8. Staying Engaged with Experiment 19 Acids And Basic Salts
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Experiment 19 Acids And Basic Salts
- 9. Balancing eBooks and Physical Books Experiment 19 Acids And Basic Salts
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Experiment 19 Acids And Basic Salts
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Experiment 19 Acids And Basic Salts
 - Setting Reading Goals Experiment 19 Acids And Basic Salts
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Experiment 19 Acids And Basic Salts

- Fact-Checking eBook Content of Experiment 19 Acids And Basic Salts
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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