



Figure 1

Crucibles are fired at high temperatures to volatilize impurities.

# Experiment 7

## Empirical Formulas

- To determine the empirical formula of two compounds by combination reactions.
- To determine the mole ratio of the decomposition products of a compound.

### OBJECTIVES

The following techniques are used in the Experimental Procedure:



### TECHNIQUES

The **empirical formula** of a compound is the simplest whole-number ratio of moles of elements in the compound. The experimental determination of the empirical formula of a compound from its elements requires three steps:

- Determine the mass of each element in the sample.
- Calculate the number of moles of each element in the sample.
- Express the ratio of the moles of each element as small whole numbers.

For example, an analysis of a sample of table salt shows that 2.75 g of sodium and 4.25 g of chlorine are present. The moles of each element are

$$2.75 \text{ g} \times \frac{\text{mol Na}}{22.99 \text{ g Na}} = 0.120 \text{ mol Na}, \quad 4.25 \text{ g} \times \frac{\text{mol Cl}}{35.45 \text{ g Cl}} = 0.120 \text{ mol Cl}$$

The mole ratio of sodium to chlorine is 0.120 to 0.120. As the empirical formula must be expressed as a ratio of small whole numbers, the whole-number ratio is 1 to 1, and the empirical formula of sodium chloride is  $\text{NaCl}$ , or simply  $\text{NaCl}$ .

The empirical formula also provides a mass ratio of the elements in the compound. The formula  $\text{NaCl}$  states that 22.99 g (1 mol) of sodium combines with 35.45 g (1 mol) of chlorine to form 58.44 g (1 mol) of sodium chloride.

The mass percentages of sodium and chlorine in sodium chloride are:

$$\frac{22.99}{22.99 + 35.45} \times 100 = 39.34\% \text{ Na}, \quad \frac{35.45}{22.99 + 35.45} \times 100 = 60.66\% \text{ Cl}$$

We can determine the empirical formula of a compound from either a **combination reaction** or a **decomposition reaction**. In the **combination reaction**, a known mass of one **reactant** and the mass of the product are measured. An example of a combination reaction is the reaction of titanium with chlorine. The initial mass of the titanium and the final mass of the titanium chloride oxide product are determined. From the difference between the masses, the mass of chlorine that reacts and, subsequently, the moles of

### INTRODUCTION



Data Analysis, A



Combination reactions: two elements combine to form a compound.

# Experiments General Chemistry 10e Lab Manual

**Roman Wölfel**



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