

Dividing Exponents

Dividing exponents is where you divide terms that involve exponents, or powers. You can divide exponents in various forms, including whole numbers, negative numbers, fractions, and decimals.

The rule for dividing exponential expressions is: $a^m \div a^n = a^{m-n}$

Same bases:	Different bases:
$5^8 \div 5^4 =$ $5^{8-4} =$ $5^4 \rightarrow$ Exponential answer $5^4 = 625 \rightarrow$ Calculated answer	$25^3 \div 5^3 =$ 25 is equal to 5^2 , replace 25 with 5^2 $(5^2)^3 \div 5^3 =$ $5^6 \div 5^3 =$ $5^{6-3} =$ $5^3 \rightarrow$ Exponential answer $5^3 = 125 \rightarrow$ Calculated answer

Division Exponent And Answer Key

**Judith A. Muschla, Gary R. Muschla, Erin
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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the K_a or K_b from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate K_{a1} and K_{a2} for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1:1
 $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$
1 mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$ and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being naoh reaction with a 2nd substance $\text{H}_2\text{C}_2\text{O}_4$ aq 2NaOH aq $\text{Na}_2\text{C}_2\text{O}_4$ $4\text{H}_2\text{O}$ to differentiate an acid from a base in a solution a ph scale is used the ph scale ranges from 0 to 14 a ph of 7 is neutral in an acid base reaction

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