

Flame Test Lab

Activity Key

Background

The heat from a laboratory burner will cause the ions of some elements to give off light. Electrons will absorb the heat energy from the flame and will “jump” to a higher energy level. When the electrons return to their original energy levels, this absorbed energy is released as light. Different elements absorb and release different amounts of energy and thus produce different kinds of light. This light may be visible or invisible, infrared or ultraviolet radiation. Analysis of this light can be used to identify elements and molecules. In today’s lab, we will observe the visible range of light emission.

Objectives

1. Perform a flame test to identify the characteristic color of metal ions.
2. Identify the metal ion in an unknown solution.
3. Calculate the energy of emitted photons.

Materials: (per lab group)

10 ml dilute solutions of the following:

Barium Chloride
Calcium Chloride
Lithium Chloride
Potassium Chloride
Sodium Chloride
Strontium Chloride
Unknown Solution

7 Wood Splints

1 Bunsen Burner

Safety Goggles & Aprons

Note:

If chloride compounds are not available, metal nitrate compounds may be substituted.

Use dilute or approximately 0.1 M solutions

Unknowns:

Number the beakers 1-12.

Add solutions as follows:

1. Strontium Chloride

2. Calcium Chloride

3. Potassium Chloride

4. Barium Chloride

5. Lithium Chloride

6. Sodium Chloride

7. Calcium Chloride

8. Strontium Chloride

9. Barium Chloride

10. Potassium Chloride

11. Sodium Chloride

12. Lithium Chloride

Procedure

1. Put on your safety goggles and an apron.
2. Obtain a wood splint and the barium chloride solution.
3. Dip the wood splint into the solution and allow it to soak for 1-2 minutes.
4. Light Bunsen Burner. USE CAUTION!
5. Remove the splint from the solution and hold the end of the splint into the flame. Do not burn the splint.
6. Observe and record the color of the flame.
7. Repeat steps 2-6 using the remaining chloride solutions.
8. Obtain an unknown solution from your teacher. Record the number of your unknown.
9. Soak a wood splint in the unknown solution.
10. Hold the splint in the flame and record the color of the flame that is produced.
11. Using your data, identify the metal ion in your unknown solution.

Flame Test Lab Answer Key

Julian Xavier Morales

A red circular graphic with a gradient, appearing as a stylized flame or a light effect, positioned to the right of the author's name.

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Gunky's Adventures Jim Reuther, 2019-08-06 A few days after the passing of his beloved wife, author Jim Reuther, better known as Gunky, discovered her extraordinary letter in a handwritten notebook titled *How to Get Along Without Me*. The notebook was a simple *How to* Guide for the tasks she had done faithfully for him until the end. But one request stood out: she challenged him to continue his writings. In *Gunky's Adventures*, Reuther features a collection of twenty-five tales, one for each letter of the alphabet, beginning with his late wife's note "Afterlife Love Letter" and "Wish," ranging from the humorous to tear-jerkers to odd happenings and surprise endings to musings on rock and roll to stories about family, friends, foes, and fails. He reflects on an array of life experiences. His first poems ever written are included under the title of "Xtraordinaire." *Silent Sentinels* Narrating a life-hearted series of alphabetic escapades, *Gunky's Adventures* offers an anthology of poems and short stories reflecting on a life wonderfully lived. **Laboratory Manual for Chemistry** Karen C. Timberlake, 1976 English

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