Name:	Date:

Student Exploration: Calorimetry Lab

Vocabulary: calorie, calorimeter, joule, specific heat capacity

Prior Knowledge Questions (Do these BEFORE using the Gizmo.)

- The Latin word color means "heat," and meter comes from the Greek word meaning "to measure." What do you think
 a calorimeter does? Measures heat energy in an object
- 2. Where have you heard the word calorie before? What do you think a calorie is? A measure of heat energy

Gizmo Warm-up

A calorimeter is an insulated container filled with a liquid, usually water. When a hot object is placed in the calorimeter, heat energy is transferred from the object to the water and the water heats up. Calorimeters can be used to find a substance's specific heat capacity. You will use the Calorimetry Lab Gizmo to determine the specific heat capacities of various substances.

- On the SIMULATION pane, select Copper. Use the slider to set its Mass to 200 g. Set the Water mass to 200 g. Check that the Water temp is set to 30.0 °C and the copper's Temp is 90 °C. Select the GRAPH tab, and click Play.
- A. What was the Final temperature of the copper and the water? 34.96 oC
- B. How much did the temperature of the copper change? -55.04 oC
- C. How much did the temperature of the water change? 4.96 oC
- 2. Specific heat capacity can be described as a substance's resistance to temperature changes. Which substance has a greater specific heat capacity, copper or water? Explain. Water, because the temperature of the water did not change that much. It would have taken more energy to raise the temperature of the water any higher than it already is.

Activity A:

Heat transfer

Get the Gizmo ready:

Click Reset.

Question: What factors determine how heat energy transfers between objects?

Predict: In the Gizmo warm-up, you saw how 200 g of 90 °C copper transfers heat to 200 g

of 30.0 °C water.

- A. How do you think increasing the water's mass would affect the final temperature? It would lower the temperature of the water a little bit
- How do you think decreasing the copper's mass would affect the final temperature? It would increase the temperature of the water a little bit
- C. How do you think increasing or decreasing the copper's initial temperature would affect the final temperature? There would be a slight raise in temperature

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