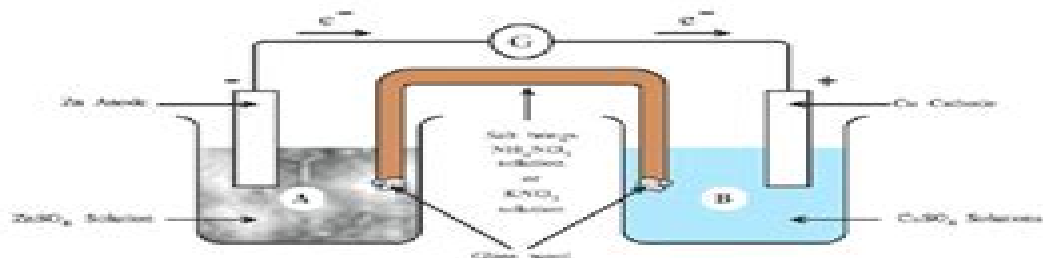


## UNIT : 8 REDOX REACTIONS & ELECTROCHEMISTRY

### Important Points

The device which transforms the chemical energy into electrical energy is called electrochemical cell. In the electrochemical cell, the energy producer in chemical reaction is transformed into electrical energy.



the salt bridge joins the two solutions and keeps the neutrality of electrical charge.

The cell in which the concentration of the ions associated with the reaction of the cell are 1 more litre<sup>-1</sup> and if any gas is associated with the reaction, then its pressure equal to 1 bar and if temperature is 298 K, is called standard cell.

The relative intensity of tendency of receiving the electron is called reduction potential  $E_{red}$  or  $E^{\circ}_{red}$  the relative intensity of electrodes for releasing electron is called oxidation potential

$E_{red}$  or  $E^{\circ}_{red}$

#### Nernst Equation



we can write the concentrations of  $Cu^{2+}$  and  $Zn^{2+}$  ions in Daniell cell, in the given electrode potential.

$$E_{cell} = E^{\circ}_{cell} - \frac{RT}{2F} \ln \left[ \frac{Zn^{2+}}{Cu^{2+}} \right]$$

#### Concentration Cell :

If the two electrodes of the electrochemical cell are same but the Concentration of the solution of the electrolyte are different, then it is called concentration cell. viz



$$\text{Cell reaction: } Cu^{2+}(C_2) \rightleftharpoons Cu^{2+}(C_1) \quad E_{cell} = E^{\circ}_{cell} - \frac{0.059}{n} \log \left[ \frac{C_1}{C_2} \right]$$

# Electrochemistry Practical Manual

**Harry Brenan Cronshaw, F.  
Hirst, William Bennett Adam**



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### **Table of Contents Electrochemisty Practical Manual**

1. Understanding the eBook Electrochemisty Practical Manual
  - The Rise of Digital Reading Electrochemisty Practical Manual
  - Advantages of eBooks Over Traditional Books
2. Identifying Electrochemisty Practical Manual
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Electrochemisty Practical Manual
  - User-Friendly Interface
4. Exploring eBook Recommendations from Electrochemisty Practical Manual
  - Personalized Recommendations
  - Electrochemisty Practical Manual User Reviews and Ratings
  - Electrochemisty Practical Manual and Bestseller Lists

5. Accessing Electrochemistry Practical Manual Free and Paid eBooks
  - Electrochemistry Practical Manual Public Domain eBooks
  - Electrochemistry Practical Manual eBook Subscription Services
  - Electrochemistry Practical Manual Budget-Friendly Options
6. Navigating Electrochemistry Practical Manual eBook Formats
  - ePub, PDF, MOBI, and More
  - Electrochemistry Practical Manual Compatibility with Devices
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7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electrochemistry Practical Manual
  - Highlighting and Note-Taking Electrochemistry Practical Manual
  - Interactive Elements Electrochemistry Practical Manual
8. Staying Engaged with Electrochemistry Practical Manual
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electrochemistry Practical Manual
9. Balancing eBooks and Physical Books Electrochemistry Practical Manual
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electrochemistry Practical Manual
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Electrochemistry Practical Manual
  - Setting Reading Goals Electrochemistry Practical Manual
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electrochemistry Practical Manual
  - Fact-Checking eBook Content of Electrochemistry Practical Manual
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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