

8 • Electron Configuration & the Periodic Table

AP QUESTION

- (a) Write the ground state electron configuration (short form or long form) for an arsenic atom.

As $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^3$
or As [Ar] $4s^2 3d^{10} 4p^3$ or [Ar] $3d^{10} 4s^2 4p^3$

- (b) Give one permissible set of four quantum numbers for each of the outermost electrons in a single As atom when it is in its ground state.

$4p$

4	1	-1	$+\frac{1}{2}$
---	---	----	----------------

$4p$

4	1	0	$+\frac{1}{2}$
---	---	---	----------------

$4p$

4	1	+1	$+\frac{1}{2}$
---	---	----	----------------

($-\frac{1}{2}$)

$4s$

4	0	0	$+\frac{1}{2}$
---	---	---	----------------

$4s$

4	0	0	$-\frac{1}{2}$
---	---	---	----------------

- (c) Is an isolated arsenic atom in the ground state paramagnetic or diamagnetic? Explain briefly.

Paramagnetic
 $4p^3 = \uparrow \uparrow \uparrow$

Half-filled orbitals = paramagnetic.

- (d) What ion would As be expected to form? As³⁻
Write the electron configuration for the arsenic ion.

As [Ar] $4s^2 3d^{10} 4p^6$ or As [Ar] $3d^{10} 4s^2 4p^6$

As gains electrons and has the same electron configuration as Kr.

3-	2-	1-	He
N	O	F	Ne
P	S	Cl	Ar
As	Se	Br	Kr

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 - The Rise of Digital Reading Electron Configuration Review Answer Key
 - Advantages of eBooks Over Traditional Books
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 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configuration Review Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Review Answer Key
 - Personalized Recommendations
 - Electron Configuration Review Answer Key User Reviews and Ratings
 - Electron Configuration Review Answer Key and Bestseller Lists
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 - Highlighting and Note-Taking Electron Configuration Review Answer Key
 - Interactive Elements Electron Configuration Review Answer Key
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configuration Review Answer Key
9. Balancing eBooks and Physical Books Electron Configuration Review Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configuration Review Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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- Setting Reading Goals Electron Configuration Review Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electron Configuration Review Answer Key
- Fact-Checking eBook Content of Electron Configuration Review Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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