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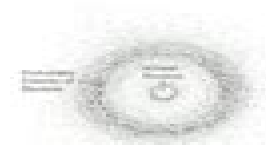
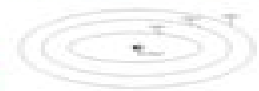
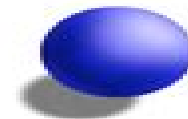
Atomic Structure & Electron Arrangement Test Review

Practice Questions:

1. The atomic number is equal to the number of **protons**.
2. The mass number is equal to the **protons+ neutrons**.
3. Electrons in the outermost energy level are known as **valence electrons** and are available to be lost, gained or shared when molecules are formed.
4. What causes an atom to be neutrally charged?
An equal number of protons and electrons
5. Describe how an atom would become negatively charged.
To be negatively charged, an atom must have gained more electrons.

Identify the major contribution of each of the following to current atomic theory (be sure to include any model that they proposed and the experiment that they did):

6. Aristotle – **everything was made from earth, wind, fire, water, and ether; alchemist**
7. Democritus – **first to propose the idea of an atom (indivisible particle)**
8. Dalton – **developed the basis for the modern atomic theory**
9. Thomson - **used the cathode ray tube to discover the first subatomic particle - the electron; developed the Plum Pudding Model**
10. Rutherford - **Gold foil experiment- expected all of the radiation to pass through and was very surprised when some of the particles were deflected - led to the discovery of the NUCLEUS (centrally located positive charge) and the concept that the atom is mostly empty space. Also discovered the proton and neutron.**
11. Bohr - **His atomic model (planetary model) had atoms built up of successive orbital shells of electrons. Electrons move between shells based on specific amounts of energy being lost or gained.**
12. Electron Cloud Model - **This model is based on quantum mechanics which states that you cannot know the speed and location of a particle. So this model shows the most probable location of the electrons.**



Electron Structure Test Review Guide

JS Bruner



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