

D. Structure of Molecules and Molecular Ions

On a separate sheet of paper, set up the following table (with eight columns) for each of the molecules/molecular ions that are assigned to you/your group. The central atom of the molecule/molecular ion is italicized.

Molecule or Molecular Ion	Lewis Structure	Valence Shell Electron Pairs	Bonding Electron Pairs	Nonbonding Electron Pairs	VSEPR Formula	Approx. Bond Angle	Geometric Shape
1. CH_4	$ \begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{H} \\ \\ \text{H} \end{array} $	4	4	0	AX_4	109.5°	tetrahedral

2. SF_6

3. H_2O

- Complete the table (as outlined above) for the following molecules/molecular ions, all of which obey the Lewis octet rule. Complete those that are assigned by your laboratory instructor.
 - H_2O^+
 - NH_3
 - NH_4^+
 - CH_3^-
 - SnH_4
 - BF_4^-
 - PO_4^{3-}
 - PF_5
 - AsH_3
 - NH_2^-
 - SF_6
 - H_2S
- Complete the table (as outlined above) for the following molecules/molecular ions, some of which obey the Lewis octet rule. Complete those that are assigned by your laboratory instructor.
 - GaI_3
 - PCl_2F_3
 - BrF_3
 - XeF_2
 - XeF_4
 - XeOF_4
 - XeOF_3
 - XeOF_4
 - SbF_6^-
 - SF_6
 - SnF_6^{2-}
 - IF_6^-
 - IF_6^+
- Complete the table (as outlined above) for the following molecules/molecular ions. No adherence to the Lewis octet rule is indicated. Complete those that are assigned by your laboratory instructor.
 - AsF_3
 - ClO_2^-
 - CF_3Cl
 - SnF_2
 - SnF_4
 - PF_6^+
 - PF_5
 - SO_4^{2-}
 - CN_2^{2-}
 - KrF_2
 - TeF_6
 - AsF_5
- Complete the table (as outlined above) for the following molecules/molecular ions. For molecules or molecular ions with two or more atoms considered as central atoms, consider each atom separately in the analysis according to Table D3.3. Complete those that are assigned by your laboratory instructor.
 - $\text{O}(\text{PCl}_2)_2$
 - H_2CCH_2
 - CH_3NH_3^+
 - Cl_2CCF_3
 - Cl_2O
 - ClCN
 - COCl_2
 - OCCCCO
 - CO_2
 - O_3
 - NO_3^-
 - $\text{TeF}_6(\text{CH}_3)_4$

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