



Dry Lab 3

Atomic and Molecular Structure

Metallic cations heated to high temperatures produce characteristic colors that appear in the starbursts.

- To view and calibrate visible line spectra
- To identify an element from its visible line spectrum
- To identify a compound from its infrared spectrum
- To predict the three-dimensional structure of molecules and molecular ions

OBJECTIVES

Visible light, as we know it, is responsible for the colors of nature—blue skies, green trees, red roses, orange-red rocks, and brown deer. Our eyes are able to sense and distinguish the subtleties and the intensities of those colors through a complex, naturally designed detection system—our eyes. The beauty of nature is a result of the interaction of sunlight with matter. Since all matter consists of atoms and molecules, it is obvious that sunlight, in some way, interacts with them to produce nature's colors.

Within molecules of compounds are electrons, vibrating bonds, and rotating atoms, all of which can absorb energy. Because every compound is different, every molecule of a given compound possesses its own unique set of electronic, vibrational, and rotational **energy states**, which are said to be **quantized**. When incident electromagnetic (EM) radiation falls on a *molecule*, the radiation absorbed (the absorbed light) is an energy equal to the difference between two energy states, placing the molecule in an excited state. The remainder of the EM radiation passes through the molecule unaffected.

Atoms of elements interact with EM radiation in much the same way, except there are no bonds to vibrate or atoms to rotate. Only electronic energy states are available for energy absorption.

EM radiation is energy as well as light, and light has wavelengths and frequencies. The relationship between the energy, E , and its wavelength, λ , and frequency, ν , is expressed by the equations:

$$E = \frac{hc}{\lambda} = h\nu \quad (\text{D3.1})$$

where h is Planck's constant, 6.63×10^{-34} J·s/photon; c equals the velocity of the EM radiation, 3.00×10^8 m/s; λ is the wavelength of the EM radiation in meters; and ν (pronounced "new") is its frequency in reciprocal seconds (s^{-1}).

EM radiation includes not only the wavelengths of the visible region (400 to 700 nm) but also those that are shorter (e.g., the ultraviolet and X-ray regions) and longer (e.g., the infrared, microwave, and radio wave regions). See Figure D3.1, page 162. From Equation D3.1, shorter wavelength EM radiation has higher energy.

When an atom or molecule absorbs EM radiation from the visible region of the spectrum that is excited from a lower to a higher energy state.

INTRODUCTION

Energy state: the amount of energy confined within an atom or molecule, which can be changed by the absorption or emission of discrete (quantized) amounts of energy

Quantized: only a definitive amount (of energy)

Wavelength: the distance between two crests of a wave

Frequency: the number of crests that pass a given point per second

Dry Lab 3 Atomic Molecular Structure Answers

J. A. Beran, Mark Lassiter



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